

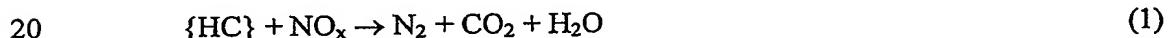
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IAP20 Rec'd PCT/PTO 09 FEB 2006

CATALYST STRUCTURE COMPRISING LEAN NO_x CATALYST
COMPOSITION

5 This invention relates to a catalyst structure for treating exhaust gas of a lean burn internal combustion engine comprising a lean NO_x catalyst (LNC) composition. In particular, the invention concerns a catalyst structure wherein the LNC composition is selected from the group consisting of: (a) silver or a silver compound supported on alumina; and (b) at least one metal selected from the group consisting of copper (Cu),
10 iron (Fe), cobalt (Co) and cerium (Ce) supported on at least one zeolite.

LNCs are sometimes also referred to in the literature as non-selective catalytic reduction (NSCR) catalysts, hydrocarbon selective catalytic reduction (HC-SCR) catalysts, lean NO_x reduction catalysts, "DeNO_x catalysts" and NO_x occluding catalysts.

15 In lean NO_x catalysis, hydrocarbons (HC) react with nitrogen oxides (NO_x), rather than oxygen (O₂), to form nitrogen (N₂), carbon dioxide (CO₂) and water (H₂O) according to reaction (1):



The competitive, non-selective reaction with oxygen is given by reaction (2):



25 A number of catalysts are known for selectively promoting the reaction (1) including platinum (Pt) on alumina (Al₂O₃), copper (Cu)-substituted zeolite such as Cu/ZSM-5 and silver (Ag) supported on Al₂O₃ (see e.g. EP 0658368).

30 Pt/Al₂O₃ is active in a relatively narrow temperature window at relatively low temperature (peak activity ~250°C). Another problem is that it is relatively unselective for N₂ formation in that it favours the formation of N₂O over N₂ at relatively low temperatures. N₂O is a potent greenhouse gas (at least 200 times more potent than CO₂) and accordingly its release into the atmosphere is undesirable.

Zeolite-based LNCs, such as Cu/ZSM5, and Ag/Al₂O₃ LNCs have wider temperature windows of activity than Pt-based LNCs and also operate at higher temperatures (peak activity of about 400-450°C and above).

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It would be desirable to develop an exhaust system for treating exhaust gas of a lean burn internal combustion engine for selectively reducing NO_x to N₂ in the presence of a suitable reductant over a relatively broad temperature window of activity without producing N₂O.

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We have investigated the known silver- and zeolite-based LNCs and we have found that their low temperature activity is improved when the reductant is treated with a partial oxidation catalyst. We believe that this step produces an oxygenated hydrocarbon feedstock. Furthermore, we believe that the activity of Ag/Al₂O₃ is suppressed by coke formation, which may be one reason why it has not found wider acceptance in the industry. We have now discovered that selective, low temperature NO_x reduction can be improved by combining known silver- and/or zeolite-based LNCs with at least one partial oxidation catalyst (POC).

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A partial oxidation catalyst (POC) is a catalyst that promotes the partial oxidation of hydrocarbons in exhaust gas of a lean-burn internal combustion engine to carbon monoxide (CO), hydrogen gas (H₂) and partially oxygenated hydrocarbon species, as opposed to complete oxidation to H₂O and CO₂.

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EP 1008379 describes a method and device for treating NO_x in an internal combustion engine exhaust system comprising a POC upstream of a NO_x trap, wherein hydrocarbons introduced into the exhaust system upstream of the POC are converted to CO and H₂ for regenerating the NO_x trap.

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EP 0761289 describes an exhaust gas cleaner and method for removing NO_x by reduction from a lean burn internal combustion engine exhaust gas. According to the disclosure, nitrogen-containing compounds such as alkyl nitrites, ammonia etc. and aldehydes are produced on e.g. Ag/Al₂O₃ which are subsequently reduced by being brought into contact with a second catalyst of: (a) titania; (b) a titanium component

supported on an inorganic oxide; (c) an inorganic oxide carrying at least one first component selected from the group consisting of copper, nickel, silver, cobalt and compounds thereof and at least one second component selected from the group consisting of titanium, zirconium and compounds thereof and a third component selected
5 from the group consisting of alkali metal elements and rare earth elements; and (d) an inorganic oxide carrying a titanium component and at least one platinum group metal. The technique of non-selective catalytic reduction is specifically distinguished in the background to the disclosure.

10 According to a first aspect, the invention provides a catalyst structure for treating exhaust gas from a lean burn internal combustion engine, which catalyst structure comprising a substrate monolith comprising a lean NO_x catalyst (LNC) composition associated with at least one partial oxidation catalyst (POC), wherein the LNC composition is selected from the group consisting of: (a) silver or a silver compound
15 supported on alumina; and (b) at least one metal selected from the group consisting of copper (Cu), iron (Fe), cobalt (Co) and cerium (Ce) supported on at least one zeolite, and wherein the at least one POC is selected from the group consisting of: (i) a bulk oxide, a bulk composite oxide or a bulk mixed oxide comprising at least one metal selected from the group consisting of manganese (Mn), iron (Fe), cerium (Ce) and praseodymium (Pr);
20 and (ii) at least one of rhodium (Rh) and palladium (Pd) disposed on at least one inorganic oxide support.

25 It is believed that where the silver is present as a compound thereof, it is present as its oxide. In use, the silver can be present in the form of its carbonate, hydroxide or nitrate by combining with components of the exhaust gas.

An advantage of the present invention is that low temperature NO_x conversion activity e.g. from 200-350°C is enhanced compared with Ag/Al₂O₃ *per se* or, for example, Cu/ZSM5 *per se*.

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A further advantage of the Ag/Al₂O₃ embodiment of the invention is that the arrangement reduces or prevents coke formation on the LNC during extended periods at low temperature, improving the overall activity of the LNC.

For the purposes of comparison, we have investigated the catalyst described in JP 2002370031 which contains Ag on mixed oxides (3-40wt.% CeO₂ on alumina) and found it to be inactive as a LNC.

5 The catalyst structure of the present invention can be arranged in a number of configurations, some of which are shown in Figure 10. The configuration can be selected according to the activity required within a desired temperature window. In one embodiment the substrate monolith comprises a physical mixture of the LNC composition and the at least one POC. In another embodiment, the substrate monolith
10 comprises a layer of the at least one POC on a layer of the LNC composition. Alternatively, the substrate monolith comprises a layer of the LNC composition on a layer of the at least one POC.

A suitable silver loading in the LNC composition of group (a) is in the range 0.5
15 to 10.0 wt.% based on the total weight of the alumina. Any form of alumina can be used, in the present invention, such as alpha-, delta- or theta-alumina. However, we have found that the best performance is obtained when using high surface area gamma-aluminas.

20 In the LNC composition, the at least one zeolite of the LNC composition can be selected from the group consisting of ZSM-5, ZSM-8, ZSM-11, ZSM-12, ZSM-20, mordenite, gamma-zeolite, beta-zeolite, silicalite, X zeolite, Y zeolite, L zeolite, erionite, USY zeolite or any mixture of two or more thereof.

25 The total amount of the at least one metal selected from the group consisting of Cu, Fe, Co and Ce in the LNC composition of group (b) is suitably in the range 0.5 to 10.0 wt.% based on the total weight of the zeolite.

Referring to the at least one POC, the oxides of group (i) can comprise at least
30 one stabiliser selected from the group consisting of: zirconium (Zr), lanthanum (La), aluminium (Al), yttrium (Y), Pr and neodymium (Nd). The at least one metal can form composite oxides or mixed oxides with the stabiliser. Suitable molar ratios of the at least one metal to the at least one stabiliser (in total) is 2:98M to 90:10M, where M is the at least one metal selected from the group consisting of Mn, Fe, Ce and Pr.

“Composite oxide” as defined herein means a largely amorphous oxide material comprising oxides of at least two elements which are not true mixed oxides consisting of the at least two elements.

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A preferable POC of group (i) for use in the invention consists of bulk CeO_2 , or a mixed oxide or composite oxide of Ce and Zr.

10 The at least one inorganic oxide support of the POC of group (ii) can comprise an oxide of Al, Ce, Zr, titanium (Ti), silicon (Si), magnesium (Mg), chromium (Cr) or a mixture, composite oxide or mixed oxide of any two or more thereof, such as a mixed oxide or a composite oxide of Ce and Zr.

15 In other embodiments, the composite oxide or mixed oxide comprises a zeolite, a non-zeolite silica-alumina, a silica-zirconia, an alumina-zirconia, an alumina-chromia, an alumina-ceria or a ceria-titania.

20 A suitable total loading of the Rh and/or Pd in the at least one POC of group (ii) is in the range 0.5 to 10.0 wt.% of Rh and/or Pd, based on the total weight of the at least one support.

25 The or each POC oxide of group (i) and/or the or each POC support of group (ii) can be doped with at least one of lanthanum (La), barium (Ba), Ce, tungsten (W), Si and Mn to improve the stability of the support to high temperature ageing, act as a lean NO_x promoter and/or improve the oxygen storage capacity of the system, the latter particularly in the case of manganese. Tungsten is a particularly suitable dopant for improving the thermal stability of group (ii) supports such as alumina or titania.

Suitable values of dopant addition are from 0.5 to 20wt%.

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Typical weight ratios of the LNC composition to the at least one POC are in the range 20:1 to 1:5, optionally from 10:1 to 1:1.

Methods of preparing the LNC compositions according to the invention are known and include wet impregnation, ion-exchange and co-precipitation. A detailed discussion of which methods will not be made here, as the skilled person is aware of such techniques.

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According to a second aspect, the invention provides an exhaust system for a lean-burn internal combustion engine comprising a catalyst structure according to the invention.

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The catalyst structure of the present invention can be configured in a number of ways, the activity of some of which are shown in the accompanying Examples and Figures. In one embodiment, some or all of the LNC composition on the substrate monolith is located downstream of the at least one POC. In another embodiment, some or all of the LNC composition is located on a separate substrate monolith disposed downstream of the substrate monolith comprising the at least one POC.

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In an alternative embodiment, some of the LNC composition on the substrate monolith is located upstream of the at least one POC. In another embodiment, some of the LNC composition is located on a separate substrate monolith disposed upstream of the substrate monolith comprising the at least one POC.

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According to a third aspect according to the invention, an apparatus comprises a lean-burn internal combustion engine including an exhaust system according to the invention.

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The exhaust system can comprise means for introducing a reductant into an exhaust gas upstream of the LNC composition. Embodiments of the exhaust system wherein at least one POC is upstream of the LNC composition can optionally include an additional means for introducing a reductant into an exhaust gas upstream of at least one POC. A reason for this is that in certain embodiments, e.g. wherein the POC is CeO₂, the POC tends to combust hydrocarbons completely at higher temperatures, hence NO_x conversion is reduced on the downstream LNC composition because of lack of reductant. A second injector between the at least one POC bed and the LNC composition bed can be used to inject hydrocarbon reductant over the LNC at temperatures which are too high

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for partial oxidation on the upstream at least one POC. In the case of CeO₂, a suitable temperature for injection of hydrocarbon between the at least one POC and downstream LNC composition is above about 350°C.

5 Alternatively, a lower loading of CeO₂, or a shorter length (or thinner "stripe") of the at least one POC, can be disposed on the substrate monolith upstream of the LNC composition to increase the amount of hydrocarbon slip or breakthrough to the downstream LNC composition.

10 The means for introducing a reductant into an exhaust gas upstream of the LNC composition can comprise at least one of: means for injecting the reductant into the exhaust gas; means for adjusting the ignition timing of at least one engine cylinder; and means for adjusting the engine air-to-fuel ratio.

15 Typically, a source of hydrocarbon reductant comprises the fuel that powers the engine.

20 The apparatus can comprise means for controlling the or each reductant introducing means. In one embodiment, the control means is arranged, when in use, to introduce the reductant into the exhaust gas when the POC is from between 200-350°C in temperature, such as in the case where the POC is CeO₂ and is located upstream of the LNC composition. This arrangement can reduce or prevent complete combustion of hydrocarbon upstream of the LNC composition, to the detriment of higher temperature NO_x reduction.

25 Generally, the control means can be arranged, when in use, to introduce the reductant into the exhaust gas when the LNC composition is above 200°C in temperature.

30 Typically, the control means can include a pre-programmed electronic control unit incorporating a processor and can form part of the engine control unit, for example. The exhaust system can include sensor means for inputting signals, such as bed temperature, tailpipe NO_x composition and NO_x content of the inlet gas, in order to

regulate hydrocarbon injection, with the overall effect being to maintain a desired NO_x conversion in the system.

The engine of the apparatus according to the invention is typically a diesel engine, optionally a heavy-duty diesel engine, but it can be any engine producing lambda >1 exhaust gas e.g. a lean-burn gasoline engine or an engine powered by compressed natural gas (CNG) or liquid petroleum gas (LPG).

A suitable substrate monolith for use in the present invention can be a flow-through monolith of metal or ceramic construction. Where ceramic, the substrate can be cordierite or silicon carbide, for example. Alternatively, the catalyst structure can include a filter substrate such as a wall-flow filter.

According to fourth aspect, the invention provides a method of selectively reducing NO_x in an exhaust gas of a lean burn internal combustion engine to N₂, which method comprising introducing a reductant into the exhaust gas and contacting the resulting mixture with a lean NO_x catalyst (LNC) composition associated with at least one partial oxidation catalyst (POC), wherein the LNC composition is selected from the group consisting of: (a) silver or a silver compound supported on alumina; and (b) at least one metal selected from the group consisting of copper (Cu), iron (Fe), cobalt (Co) and cerium (Ce) supported on at least one zeolite, and wherein the at least one POC is selected from the group consisting of: (i) a bulk oxide, a bulk composite oxide or a bulk mixed oxide comprising at least one metal selected from the group consisting of manganese (Mn), iron (Fe), cerium (Ce) and praseodymium (Pr); and (ii) at least one of rhodium (Rh) and palladium (Pd) disposed on at least one inorganic oxide support.

If desired, the catalyst structure of the present invention can be combined with one or more additional LNCs to yield an exhaust system with yet further improved activity across a desired temperature window. In practical applications it may be desirable to include an oxidation catalyst to remove excess reductant which slips the LNC composition.

In order that the invention may be more fully understood, the following Examples are provided by way of illustration only and with reference to the accompanying drawings, in which:

5 Figure 1 is a graph comparing %NO_x conversion between steady state and transient modes for Comparative Example 1 as a function of temperature;

10 Figure 2 is a graph showing %NO_x conversion over the Catalyst of Example 1 compared with the Catalyst of Comparative Example 2 in the steady state mode as a function of temperature;

15 Figure 3 is a graph showing the outlet NO_x concentration as a function of time at 200°C, 250°C and 300°C over 2%Ag/Al₂O₃ and the 2Ag/Al₂O₃-CeO₂ mixture in the steady mode;

Figure 4 is a graph showing the effect of ramp down (steady state 15 minutes) and ramp up (steady state 15 minutes) on NO_x conversion of 2Ag/Al₂O₃-CeO₂ (4:1);

20 Figure 5 is a graph showing the effect of ramp down (steady state 15 minutes) and ramp up (steady state 15 minutes) on HC conversion of 2Ag/Al₂O₃-CeO₂ (4:1);

Figure 6 is a graph comparing the effect of ageing on Comparative Example 1 and Example 1 catalysts on NO_x conversion;

25 Figure 7 is a graph showing %NO_x conversion as a function of temperature for 5Cu/ZSM5 catalyst structures compared with 5Cu/ZSM5 catalyst *per se*;

Figure 8 is a graph showing %NO_x conversion as a function of temperature for 2Ag/Al₂O₃ catalyst structures compared with 2Ag/Al₂O₃ catalyst *per se*;

30 Figure 9 is a graph showing %NO_x conversion as a function of temperature for a catalyst according to JP 2002370031 compared with 2Ag/Al₂O₃ and 5Cu/ZSM5 catalyst structures according to the invention; and

Figure 10A-F includes a series of schematic arrangements of catalyst structures according to the present invention.

Referring to Figure 10, Figure 10A is a schematic diagram representing an 5 exhaust line including a can comprising a catalyst structure according to the invention comprising a ceramic flow-through substrate monolith disposed between two diffusers. "HC" represents a schematic injection point for hydrocarbon reductant. The upstream end zone (as shown by the direction of flow arrows) of the substrate monolith comprises an upper layer of a POC (such as bulk CeO₂) and an underlayer of LNC composition, 10 such as Ag/A1₂O₃.

In an alternative configuration of the Figure 10A embodiment shown in Figure 10B, the underlayer in the upstream zone is the POC, whereas the overlayer is the LNC.

15 Figure 10C shows a further alternative to the arrangement shown in Figure 10A, wherein the POC is disposed in an overlayer on the LNC on a first substrate monolith and a second substrate monolith coated with LNC is disposed downstream. An additional injector is disposed between the first and second substrate monoliths and a control means (not shown) is arranged to inject hydrocarbon between the two substrate 20 monoliths e.g. continuously when the downstream LNC is above e.g. 200°C (when a Ag/A1₂O₃ or Cu/ZSM5 catalyst is used) and to inject hydrocarbon upstream of the upstream substrate monolith at, for example, 200-350°C in the case of CeO₂ or Ce-containing mixed- or composite-oxides. This arrangement enables hydrocarbon to contact the downstream LNC at higher temperatures in order to promote NO_x reduction, 25 which hydrocarbon would otherwise be combusted on the upstream POC if the exhaust system comprised only an hydrocarbon injector upstream of the POC. Optionally, the Figure 10C arrangement can also include an additional injector between the upstream and downstream zones for similar reasons. The first and second substrate monoliths of Figure 10C can be inserted into the same can or separate cans in the exhaust line.

30 Alternatively, the LNC can be disposed in a zone on the front (upstream) end of the substrate monolith, and the LNC can be arranged in an overlayer on the POC as shown in Figure 10D.

The length of the zones in each of the embodiments represented in Figure 10 can be selected according to the desired activity of the zones, the active catalyst used e.g. to enable thrifting of catalyst etc. In the arrangement shown in Figure 10E, for example, a short zone or "stripe" of POC is used on the upstream end to limit or "trim" HC combustion in the catalyst structure at higher temperature.

The Figure 10F embodiment shows an arrangement wherein the LNC and POC are coated on the catalyst substrate as a physical mixture or the entire length of the substrate monolith includes an overlayer of POC and an underlayer of LNC or vice versa.

Comparative Example 1

2 wt%Ag/Al₂O₃Catalyst Preparation

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An impregnated Ag catalyst at 2wt% Ag based on the weight of the Al₂O₃ support (referred to hereinafter as 2Ag/Al₂O₃) was made using silver nitrate by a wet impregnation method. The silver nitrate was dissolved in the appropriate amount of water measured from the pore volume of an Al₂O₃ support. The solution was then added to the Al₂O₃ with mixing. After drying overnight, the sample was calcined in air at 500°C for 2 hours. The resulting catalyst was pelleted and particles in the size range 250< d <355 μm were selected.

Comparative Example 2

Mixed 4:1 2 wt%Ag/Al₂O₃:cordierite Catalyst Preparation

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The pelleted catalyst particles in the size range 250< d <355 μm prepared according to Comparative Example 1 were physically mixed with pelleted particles of cordierite of the same size in a weight ratio of 4:1 2Ag/Al₂O₃:cordierite.

Example 1**4:1 2 wt%Ag/Al₂O₃:CeO₂ Catalyst Preparation**

5 The pelleted catalyst particles in the size range 250< d <355 μm prepared according to Comparative Example 1 were physically mixed with pelleted particles of bulk ceria of the same size in a weight ratio of 4:1 2Ag/Al₂O₃:ceria.

Example 2**10 Coated substrate including 2 wt%Ag/Al₂O₃**

Three cores from ceramic flow-through monolith substrates were prepared, each 2.54cm (1 inch) in diameter and 7.62 cm (3 inches) in length. A first core was coated with a washcoat containing the powder form of the catalyst of Comparative Example 1. 15 The Ag/Al₂O₃ loading was 2.5g per cubic inch (152.6g per litre). A second core was coated with a first layer of CeO₂ powder and a second layer (overlayer) of the catalyst of Comparative Example 1 at the same loading as the first core. The catalyst loading of the 2Ag/Al₂O₃:CeO₂ was 4:1. In the third core, one end of the core was coated with a first layer of the same amount of CeO₂ powder as the second core to a depth of 2.54 cm (one 20 inch), and a second layer (overlayer) of the catalyst of Comparative Example 1 at the same loading as the first core. The catalyst loading of 2Ag/Al₂O₃:CeO₂ on the whole of the third core was 4:1.

Example 3**25 Coated substrate including 5 wt%Cu/ZSM5**

Example 2 was repeated except in that a 5wt% Cu/ZSM5 catalyst prepared by wet impregnation was used instead of the powdered 2Ag/Al₂O₃ catalyst.

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Comparative Example 3**Coated Substrate including catalyst of JP 2002370031**

A ceramic flow-through monolith substrate core 2.54cm (1 inch) in diameter and 7.62 cm (3 inches) in length was coated with alumina powder at a loading of 2.5 g per

cubic inch (152.6g per litre) and the coated core was dried. The washcoated alumina was then impregnated with a mixture of silver and cerium salts to achieve a final catalyst loading of 2wt% silver and 10wt% cerium, and the resulting piece was dried and calcined. As far as the abstract of this patent publication can be understood, the above 5 arrangement falls within its teaching.

Example 4

Activity Measurements

10 The activity of the pelleted catalysts to reduce NO to N₂ using MK1 (diesel fuel) as the reductant was measured in a simulated catalyst activity test (SCAT) gas rig species. The simulated exhaust gas composition used in the test was as follows: NO 500ppm, C1 2250ppm (MK1 diesel fuel), CO 200ppm, O₂ 9%, H₂O 4.5%, CO₂ 4.5%, N₂ balance (C1:NO_x 4.5:1). The data was collected in one of two modes:

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- (i) steady state conditions with 15 minutes dwell at each temperature, (data points collected at the end of 15 minutes); or
- (ii) transient ramp up at 10°C per minute temperature from 200°C to 500°C.

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The coated cores were tested only under steady state conditions (data collection as above) in the following simulated exhaust gas composition: NO 500ppm, C1 2250ppm (US06 diesel fuel), CO 200ppm, O₂ 8%, H₂O 4.5%, CO₂ 4.5%, N₂ balance (C1:NO_x 4.5:1).

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Comparative Example 4

Results of Activity Measurements on Catalyst of Comparative Example 1

Figure 1 shows a graph comparing %NO_x conversion between steady state and 30 transient modes for the catalyst of Comparative Example 1 as a function of temperature. It can be seen that the sample is very active for NO_x conversion between 300°C and 550°C (~60% conversion) in the transient mode, but it is severely deactivated in the steady state mode compared with the transient mode. We believe that this deactivation is caused by carbon species ("coke") deposited on the catalyst surface after exposure to the

gas mixture for long periods at low temperature. Such conditions can be encountered in the exhaust of a diesel engine. Reasons for the above conclusion include our finding that catalyst deactivation is less severe at a C1:NO_x ratio of 3:1 and activity is restored by heating the deactivated catalyst in the reaction mixture at 550°C (results not shown).

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Example 5

Results of Activity Measurements on Catalyst of Example 1

Figure 2 shows the %NO_x conversion activity over the 2Ag/Al₂O₃-CeO₂ catalyst of Example 1 compared with the 2Ag/Al₂O₃ catalyst of Comparative Example 2 in the steady state mode as a function of temperature. The results show that the addition of particulate CeO₂ to the 2Ag/Al₂O₃ catalyst greatly enhances its activity between 200°C-350°C. We believe that this results from generation of partial oxidation products from the MK1 hydrocarbon in the low temperature region which promote NO_x reduction on the 2Ag/Al₂O₃ LNC. The control sample mixed with cordierite has no significant activity in this temperature region. Higher temperature activity is lower with the 2Ag/Al₂O₃-CeO₂ mixture due to the non-selective HC oxidation over the CeO₂, but this is dependent on the configuration of the catalyst structure as shown in Example 9 and Figure 8.

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Figure 3 shows the outlet NO_x concentration (ppm) as a function of time at 200°C, 250°C and 300°C over a 2Ag/Al₂O₃ catalyst of Comparative Example 1 and the 2Ag/Al₂O₃-CeO₂ catalyst of Example 1 in the steady mode. The detectors were switched to measure inlet NO_x concentration during periods of temperature adjustment. The 2Ag/Al₂O₃-CeO₂ catalyst shows substantially no decay in activity during the 15 minute steady state data collection period at 250°C and 300°C compared with the 2Ag/Al₂O₃ catalyst. We believe that this is due to a combination of: the absence of any coke deposition on the 2Ag/Al₂O₃-CeO₂ catalyst; and the feed stock containing reformed i.e. partially oxidised MK1 hydrocarbon products that are particularly effective for promoting NO_x reduction. It can be seen that at 200°C, NO_x conversion tails off over the 2Ag/Al₂O₃ catalyst, possibly as a result of coking of the catalyst.

Example 6Results of Activity Measurements: Steady State Ramp-Up and Ramp-Down on Catalyst of Example 1

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The activity of the 2Ag/Al₂O₃-CeO₂ catalyst under steady state conditions from 500°C to 200°C and then from 200°C to 500°C (C1:NO_x=4.5:1) was measured and the results are shown in Figures 4 and 5 respectively, following the procedure illustrated in Figure 3. The NO_x and HC conversion are similar between ramp down and ramp up and 10 no catalyst deactivation was observed. The effect of lowering the C1:NO_x ratio to 3:1 is also shown. In contrast to previous results with 2Ag/Al₂O₃ alone, the NO_x conversion is lower at low temperature at a C1:NO_x ratio of 3:1. The results are summarised in Table 1.

15 **Table 1: Effect of CeO₂ addition on 2Ag/Al₂O₃ lean NO_x activity.**

Catalyst	NO _x conversion			HC conversion		
	200°C	250°C	300°C	200°C	250°C	300°C
2Ag/Al ₂ O ₃ +corderite	8	7	13	--	--	--
2Ag/Al ₂ O ₃ - CeO ₂	12	41	37	1	43	56
2Ag/Al ₂ O ₃ - CeO ₂ *	14	24	28	19	56	68

*C1:NO_x=3:1Example 7Catalyst Ageing

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The 2Ag/Al₂O₃ and 2Ag/Al₂O₃-CeO₂ catalysts were lean hydrothermally aged in a 10% H₂O and air mixture at 700°C for 16 hours and the steady state activity of the resulting catalysts are shown in Fig 6. It can be seen that even after the high temperature ageing, the presence of CeO₂ is beneficial to the low temperature activity.

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Example 8**Results of Activity Measurements on Catalyst Cores of Example 3**

The results of steady state analysis tests performed on the catalyst cores of Example 3, wherein the third core was oriented with the CeO₂ zone on the upstream side are shown in Figure 7. It can be seen that low temperature NO_x conversion in the 5Cu/ZSM5 core is improved for the catalyst structures containing CeO₂ relative to 5Cu/ZSM5 *per se*. We believe that this results from the supply of partially oxidised US06 hydrocarbon products to the LNC at low temperature.

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Example 9**Results of Activity Measurements on Catalyst Cores of Example 3**

The results of steady state analysis tests performed on the catalyst cores of Example 2 are shown in Figure 7. Two configurations of the third core were tested: the CeO₂ zone on the upstream end; and the CeO₂ zone on the downstream end. It can be seen that low temperature NO_x conversion in the 2Ag/Al₂O₃ core is improved for the catalyst structures containing CeO₂ relative to 2Ag/Al₂O₃ *per se*. We believe that this results from a combination of partial oxidation products in the feedstock contacting the LNC and reduced coking of the LNC. Higher temperature NO_x conversion is suppressed in the configurations wherein the CeO₂ is present in the upstream zone of the core relative to 2Ag/Al₂O₃ *per se*, because the CeO₂ favours complete combustion of the hydrocarbon instead of partial oxidation at higher temperatures. However, this can be retained with improved low temperature NO_x conversion relative to 2Ag/Al₂O₃ *per se* when the mixed 2Ag/Al₂O₃-CeO₂ zone is on the downstream end of the core with 2Ag/Al₂O₃ *per se* on the upstream end zone.

Comparative Example 4**Results of Activity Measurements on Catalyst of Comparative Example 3**

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The activity of the catalyst of JP 2002370031 tested under steady state conditions is shown in Figure 9, compared to the activity of comparable catalyst arrangements of the 5Cu/ZSM5 and 2Ag/Al₂O₃ embodiments according to the invention.